

which do not collide with the surface. In a magnetic field parallel to the surface, the formula for the magnetic moment, for any kind of reflection from the surface, coincides with formula (14) of the work by Kosevich and Lifshitz.⁴

Consequently, if one investigates $\chi_Q(H)$ in a film with $D < l$, then, by determining the successive values of the magnetic field H for which there are sharp increases in the quantum oscillations with a given period, for all possible directions of H , we determine d and S_{ext} for each of the surfaces individually (we mean here the closed surfaces into which $\epsilon(\mathbf{p}) = \epsilon_0$ is resolved), without resorting to harmonic analysis. (The applicability of the above arguments to this case is obvious.)

The values of $d = D$ in a parallel field make it possible to determine

$$p_x^{\text{max}} = |eHD|/2c$$

for the central cross-section of the above surfaces, at all directions x in the plane of the film,

i.e., to construct directly all the cross-sections $\epsilon(\mathbf{p}) = \epsilon_0$ and $p_y = 0$. A set of monocrystallographic films with a varying orientations of the crystallographic axes relative to the surface would make it possible to plot $\epsilon(\mathbf{p}) = \epsilon_0$.

The values of S_{ext} , as already indicated, also make it possible to synthesize the Fermi surface by the Lifshitz-Pogorelov method.

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CALORIMETRIC DETERMINATION OF THE HALF-LIFE OF Ra^{226}

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IN the determination of the half-lives of long-lived isotopes by the calorimetric method, one measures the thermal effect Q of the radioactivity of a known weight p of a given isotope. The half-life T is determined from the relation $T = (\ln 2) \cdot pN_0\epsilon/AQ$, where N_0 is Avogadro's number, A the atomic weight of the isotope, and ϵ the energy liberated by the isotope in the calorimeter in a single decay event. The decay constants of Pu^{239} (Ref. 1) and Ac^{227} (Ref. 2) were measured relatively recently by such a calorimetric method.

Many investigators have used this method to determine T for Ra^{226} . They measured the thermal effect of radium, either free of its short-lived derivatives³ or in equilibrium with these derivatives.^{4,5} Unfortunately, the Ra contents in the measured compounds was determined in these in-

vestigations indirectly, usually on the basis of γ measurements. In addition, for the most part, the accuracy of the calorimetric measurements themselves was insufficient (1-4%), as was the accuracy with which ϵ was calculated for the elements of the Ra^{226} series.

We made precise calorimetric measurements of three equilibrium compounds of radium, subjected to 10 (Ra I, Ra II) and 15 (Ra III) additional crystallizations to eliminate possible contaminations. The purity of these compounds was controlled by means of a spectral method, which detected the presence of only insignificant impurities of certain elements of the second group of the periodic table.⁶ The total contamination α did not exceed, in all three cases, 0.2% (by weight).

The results of direct weighing of radium bromide before sealing the compounds, and the Ra^{226} content in these compounds, are given in the table. The table lists also the results of the calorimetric measurements performed with the aid of a double static calorimeter.⁹ The last column of the table gives the determined values of $Q_{\alpha+\beta}/p$ — the thermal effect of α - and β -radiation per gram of radium. In the calculation of these quantities, account is taken of the absorption of γ -rays in the compound itself (self absorption), in the protective container, in the glass of the ampoules, and in the walls of the calorimetric cylinder itself.